

Atoms, ions,
Molecules

$\xleftarrow{\text{MSE}}$
 $\xrightarrow{6.022 \times 10^{23}}$

Moles of an
Element

$\xleftrightarrow{\text{Molar Mass}}$

GRAMS

Conversions between Moles, Atoms & m.m

Atoms, ions,
Molecules

$\xleftarrow{\text{MSE}}$
 $\xrightarrow{6.022 \times 10^{23}}$

Moles of a
Compound

$\xleftrightarrow{\text{Molar Mass}}$

GRAMS

6.022×10^{23} = Avogadro's #

Converting from Grams of substance A $\xrightarrow{\quad}$ to Mass of substance B

Grams of A $\xrightarrow{\text{Molar Mass}}$ Moles of A $\xrightarrow{\text{Mole to Mole factor } \frac{B}{A}}$ Moles of B $\xrightarrow{\text{Molar Mass of factor B}}$ Grams of B